



Catholic Junior College

JC 2 Preliminary Examinations

Higher 2

CANDIDATE
NAME

CLASS

CHEMISTRY

Paper 4 Practical

9729/04

Thursday 16 August 2018

2 hours 30 minutes

Candidates answer on the Question Paper

READ THESE INSTRUCTIONS FIRST

Write your class and name on all the work you hand in.

Give details of the practical shift and laboratory, where appropriate, in the boxes provided.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

Answer **all** questions in the spaces provided on the Question Paper.

The use of an approved scientific calculator is expected, where appropriate.

You may lose marks if you do not show you working or if you do not use appropriate units.

Qualitative Analysis Notes are printed on pages 14 and 15.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

Shift
Laboratory

For Examiner's Use	
1	17
2	18
3	20
Total	55

This document consists of **15** printed pages and **1** blank page.

- 1 **Z** is an oxidising agent. You are required to use a solution of iodine in potassium iodide to determine the mass of **Z** needed to liberate 1 mole of iodine from potassium iodide.

FA3 is an aqueous solution of iodine, I_2 , in potassium iodide, KI.

FA4 is a $0.100 \text{ mol dm}^{-3}$ aqueous sodium thiosulfate, $Na_2S_2O_3 \cdot 5H_2O$.

FA5 is an aqueous solution containing 2.00 g dm^{-3} **Z**, an oxidising agent.

You are to follow the following procedure and record your results as instructed.

Titration 1

- 1 By using a pipette, transfer 25.0 cm^3 of **FA3**, an aqueous solution of iodine in potassium iodide, into a conical flask.
- 2 Titrate the iodine present with **FA4**, $0.100 \text{ mol dm}^{-3}$ aqueous sodium thiosulfate, from a burette until the solution becomes colourless.
- 3 Repeat the titration as many times as you think necessary to obtain accurate results.
- 4 Record your results of titration in a suitable tabulated form below.

Titration of iodine in **FA3** with **FA4**

Summary

..... cm^3 of **FA3** required cm^3 of **FA4** for titration.

Show which results you used to obtain this value of the volume of **FA4** by placing a tick ($\surd$) under the readings used.

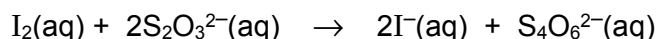
[5]

- (a) Calculate how many moles of sodium thiosulfate were contained in the volume of **FA4** found necessary to react with 25.0 cm^3 of **FA3**.

Moles of $Na_2S_2O_3 = \dots\dots\dots$

[1]

- (b) Calculate how many moles of I_2 present in 25.0 cm^3 of **FA3** using the following equation for the reaction between iodine and thiosulfate ions.



Moles of I_2 =
[1]

Titration 2

- 1 By using a pipette, transfer 25.0 cm^3 of **FA3**, an aqueous solution of iodine in potassium iodide, into a conical flask and then using a second pipette, add 25.0 cm^3 of **FA5**, an aqueous solution containing 2.00 g dm^{-3} of the oxidising agent, **Z**.
- 2 Titrate the total iodine present with **FA4**, $0.100 \text{ mol dm}^{-3}$ aqueous sodium thiosulfate, from a burette as before until the solution becomes colourless.
- 3 Repeat the titration as many times as you think necessary to obtain accurate results.
- 4 Record your results of titration in a suitable tabulated form below.

Titration of total iodine now in **FA3** with **FA4**

Summary

..... cm^3 of **FA3** to which 25.0 cm^3 of **FA5** had been added required cm^3 of **FA4** for titration.

Show which results you used to obtain this value of the volume of **FA4** by placing a tick (✓) under the readings used.

[3]

[Turn over

- (c) Calculate how many moles of sodium thiosulfate were contained in the volume of **FA4** found necessary to react with the total iodine in the flask after 25.0 cm³ of **FA3** and 25.0 cm³ of **FA5** had been mixed.

Moles of Na₂S₂O₃ =
[1]

- (d) Calculate how many moles of I₂ present in the flask after 25.0 cm³ of **FA3** and 25.0 cm³ of **FA5** had been mixed.

Moles of I₂ =
[1]

- (e) Calculate how many moles of I₂ were liberated from potassium iodide in the flask by 25.0 cm³ of **FA5**.

Moles of I₂ =
[2]

- (f) Calculate the mass of **Z** present in 25.0 cm³ of **FA5**.

Mass of **Z** =
[1]

- (g) Hence, from (e) and (f), calculate the mass of **Z** needed to liberate 1 mole of I₂.

Mass of **Z** =
[2]

[Total: 17]

[Turn over

- 2 You are required to determine the basicity of an acid by carrying out some experiments with aqueous sodium hydroxide.

FA1 is a 1.00 mol dm^{-3} aqueous sodium hydroxide, NaOH.

FA2 is a 1.00 mol dm^{-3} aqueous acid which is either monobasic with the formula **HA**, or dibasic with the formula **H₂A**.

By using the results of the following experiments, you are to going to determine the basicity of **FA2**.

You are to follow the following procedure and record your results as instructed.

Experiment 1

- 1 By using a measuring cylinder, pour 25 cm^3 of **FA1** into the Styrofoam cup provided and record the steady temperature. Rinse the thermometer with deionised water.
- 2 Use another 50 cm^3 measuring cylinder to measure 25 cm^3 of **FA2** and record its temperature.
- 3 Carefully add **FA2** to **FA1**, stir with the thermometer and record the highest temperature reached.
- 4 Record your results in the table provided below.

	Experiment 1	Experiment 2	Experiment 3
Initial temperature of FA1 /°C			
Initial temperature of FA2 /°C			
Highest temperature /°C			
Mean initial temperature /°C			
Temperature rise, ΔT /°C			

Experiment 2

Thoroughly rinse out the Styrofoam cup with deionised water and repeat the experiment using 25 cm^3 of **FA1** and 50 cm^3 of **FA2**. Record your results in the table above.

Experiment 3

Thoroughly rinse out the Styrofoam cup with deionised water and repeat the experiment using 50 cm^3 of **FA1** and 25 cm^3 of **FA2**. Record your results in the table above.

- (a) Complete the table of results by calculating the mean initial temperature of **FA1** and **FA2** for each of the three experiments. When calculating mean initial temperatures for **Experiments 2** and **3**, ignore the fact that the volumes mixed are unequal.

Fill in the corresponding rises in temperatures, ΔT in the table provided above.

[5]

(b) Use your results to deduce the basicity of the acid.

(c) Write a full equation for the reaction between **FA1** and **FA2**. Do not omit spectator ions. [2]

..... [1]

(d) Calculate the amount of heat evolved in **Experiment 1**.
Assume 4.3 J are required to raise the temperature of 1 cm³ of solution by 1.0 °C.

Heat evolved = [1]

(e) Using your answers to (c) and (d), calculate the enthalpy of neutralisation, ΔH_n , per mole of water formed.

ΔH_n = [1]

[Turn over]

Assessment of Planning Skills

You are provided with 150 cm³ of each of three solutions, **FA6**, **FA7** and **FA8**. You are to identify which of these solutions correspond to the following solutions:

1 mol dm⁻³ sodium hydroxide, NaOH

2 mol dm⁻³ hydrochloric acid, HCl

1 mol dm⁻³ hydrochloric acid, HCl

You are also provided with the following apparatus:

A thermometer

A plastic cup

2 measuring cylinders

You are to plan a simple experiment which will allow an identification of the three solutions to be made, using only the apparatus and materials listed above and deionised water. (Note that you are **not** provided with any pH indicator).

You are **not** required to carry out the plan but you have to present your results in an appropriate manner based on your plan so that the identification can be made.

PLAN

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[5]

RESULTS

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[3]

[Total: 18]

[Turn over

3 You are to determine the identity of **FA9**.

Carry out the following experiments on the solid **FA9** which contains the **sodium cation** and **one anion** from the following list: CO_3^{2-} , Cl^- , Br^- , I^- , NO_3^- , NO_2^- , SO_4^{2-} and SO_3^{2-} .

In all tests, reagents should be added gradually until no further change is observed, with shaking after each addition.

Record your observations you make from them in the spaces provided.

Your answer should include

- (i) details of colour change and precipitate formed, if any;
- (ii) the identity of gases evolved, if any, and details of the test used to identify the gas.

You should indicate clearly at what stage in a test a change occurs.

No additional or confirmatory tests for ions present should be attempted.

Test	Observation
(a) Place one third of the solid FA9 in a test-tube and heat strongly until the solid melts and a gas is given off. Identify the gas evolved. Continue the strong heating for 2–3 minutes then leave the tube to cool and retain the residue for test (d). (Melt a second sample of FA9 and heat for 2–3 minutes. Leave this to cool and retain for test (h).	
(b) Using a dropper, place 1 cm depth of aqueous of sodium hydroxide into a test-tube and tip into it the remaining solid FA9. Warm gently and retain for test (c).	
(c) Cool the test-tube and solution from (b). Add a half spatula full of aluminium powder and warm gently.	

Test	Observation
(d) Dissolve the residue from (a) in deionised water about 3 cm depth and divide the solution into three parts. Use these tests for (e) to (g).	
(e) To one part of the solution from (d), add aqueous potassium iodide followed dilute sulfuric acid.	
(f) To the second part of the solution from (d), add aqueous potassium manganate(VII) followed by dilute sulfuric acid.	
(g) To the third part of the solution from (d), add aqueous sodium hydroxide and a half spatula full of aluminium powder. Warm gently.	
(h) Take the second sample of solid, heated FA9 that you prepared in (a). Add to dilute sulfuric acid and warm gently.	

[9]

What anion is present in **FA9** before it is heated?

What anion is present in **FA9** after it is heated?

In test (e), the anion is behaving as

In test (f), the anion is behaving as

[4]

Hence, write a balanced chemical equation for the thermal decomposition of **FA9**.

.....

[1]

[Turn over]

You are provided with a set of test-tubes and the following reagents:

You are required to devise and perform an experiment which will enable you to identify the cation present in each of the solutions **FA10**, **FA11** and **FA12**. You may only use any of the reagents provided above and **no** other reagent. **You should aim to use the minimum number of reactions**, where possible.

This space below is for you to plan your steps which should then be recorded, together with the observations you make in the table in the next page.

PLAN

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[2]

In the table below, record all of your observations, and the identity of a cation as soon as this is established.

Reagent added	Observations and identities of cations		
	FA10	FA11	FA12

[3]

Deduction

Solution **FA10** contains the cation

Solution **FA11** contains the cation

Solution **FA12** contains the cation

[1]

[Total: 20]

[Turn over]

Qualitative Analysis Notes

[ppt. = precipitate]

(a) Reactions of aqueous cations

cation	reaction with	
	NaOH(aq)	NH ₃ (aq)
aluminium, Al ³⁺ (aq)	white ppt. soluble in excess	white ppt. insoluble in excess
ammonium, NH ₄ ⁺ (aq)	ammonia produced on heating	–
barium, Ba ²⁺ (aq)	no ppt. (if reagents are pure)	no ppt.
calcium, Ca ²⁺ (aq)	white ppt. with high [Ca ²⁺ (aq)]	no ppt.
chromium(III), Cr ³⁺ (aq)	grey-green ppt. soluble in excess giving dark green solution	grey-green ppt. insoluble in excess
copper(II), Cu ²⁺ (aq)	pale blue ppt. insoluble in excess	blue ppt. soluble in excess giving dark blue solution
iron(II), Fe ²⁺ (aq)	green ppt., turning brown on contact with air insoluble in excess	green ppt. turning brown on contact with air insoluble in excess
iron(III), Fe ³⁺ (aq)	red-brown ppt. insoluble in excess	red-brown ppt. insoluble in excess
magnesium, Mg ²⁺ (aq)	white ppt. insoluble in excess	white ppt. insoluble in excess
manganese(II), Mn ²⁺ (aq)	off-white ppt. rapidly turning brown on contact with air insoluble in excess	off-white ppt. rapidly turning brown on contact with air insoluble in excess
zinc, Zn ²⁺ (aq)	white ppt. soluble in excess	white ppt. soluble in excess

(b) Reactions of anions

<i>anion</i>	<i>reaction</i>
carbonate, CO_3^{2-}	CO_2 liberated by dilute acids
chloride, $\text{Cl}^-(\text{aq})$	gives white ppt. with $\text{Ag}^+(\text{aq})$ (soluble in $\text{NH}_3(\text{aq})$)
bromide, $\text{Br}^-(\text{aq})$	gives pale cream ppt. with $\text{Ag}^+(\text{aq})$ (partially soluble in $\text{NH}_3(\text{aq})$)
iodide, $\text{I}^-(\text{aq})$	gives yellow ppt. with $\text{Ag}^+(\text{aq})$ (insoluble in $\text{NH}_3(\text{aq})$)
nitrate, $\text{NO}_3^-(\text{aq})$	NH_3 liberated on heating with $\text{OH}^-(\text{aq})$ and Al foil
nitrite, $\text{NO}_2^-(\text{aq})$	NH_3 liberated on heating with $\text{OH}^-(\text{aq})$ and Al foil; NO liberated by dilute acids (colourless $\text{NO} \rightarrow$ (pale) brown NO_2 in air)
sulfate, $\text{SO}_4^{2-}(\text{aq})$	gives white ppt. with $\text{Ba}^{2+}(\text{aq})$ (insoluble in excess dilute strong acids)
sulfite, $\text{SO}_3^{2-}(\text{aq})$	SO_2 liberated with dilute acids; gives white ppt. with $\text{Ba}^{2+}(\text{aq})$ (soluble in dilute strong acids)

(c) Tests for gases

<i>gas</i>	<i>test and test result</i>
ammonia, NH_3	turns damp red litmus paper blue
carbon dioxide, CO_2	gives a white ppt. with limewater (ppt. dissolves with excess CO_2)
chlorine, Cl_2	bleaches damp litmus paper
hydrogen, H_2	“pops” with a lighted splint
oxygen, O_2	relights a glowing splint
sulfur dioxide, SO_2	turns aqueous acidified potassium manganate(VII) from purple to colourless

(d) Colour of halogens

<i>halogen</i>	<i>colour of element</i>	<i>colour in aqueous solution</i>	<i>colour in hexane</i>
chlorine, Cl_2	greenish yellow gas	pale yellow	pale yellow
bromine, Br_2	reddish brown gas / liquid	orange	orange-red
iodine, I_2	black solid / purple gas	brown	purple

[Turn over

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